

Lanthanum-Doped Ceria Nanocomposite: A Highly Stable Monolithic Catalyst for Direct Synthesis of Dimethyl Carbonate

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Abstract: That dimethyl carbonate is directly synthesized from methanol and carbon dioxide is an effective and environmental approach to solve the greenhouse effect. For the sake of solving the problems of low DMC productivity and poor catalysts stability in presence of the formed water. Here we design and prepare a series of spherical La-doped ceria nanoparticles ($\text{Ce}_{1-x}\text{La}_x\text{O}_8$ nanocomposites, $x=0.00, 0.05, 0.10, 0.15$, and 0.20) via a co-precipitation method. These $\text{Ce}_{1-x}\text{La}_x\text{O}_8$ composites are ground into slurry by ball milled and then coated on cordierite honeycomb ceramics to obtain $\text{Ce}_{1-x}\text{La}_x\text{O}_8$ monolithic catalysts. These $\text{Ce}_{1-x}\text{La}_x\text{O}_8$ composites are characterized extensively by TEM, XRD, Raman spectroscopy, N_2 adsorption-desorption isotherms, H_2 -TPR and XPS. The characterization results show that the $\text{Ce}_{1-x}\text{La}_x\text{O}_8$ composites nanoparticles possesses richer surface oxygen vacancies, higher BET surface area and smaller particle size than that of pure CeO_2 nanoparticle. Besides, catalytic activity test shows these $\text{Ce}_{1-x}\text{La}_x\text{O}_8$ monolithic catalysts exhibit better catalytic performance than that of pure CeO_2 nanoparticles. Among them, $\text{Ce}_{0.95}\text{La}_{0.05}\text{O}_8$ monolithic catalyst exhibits the highest CH_3OH conversion and DMC yield, which is in good line with the oxygen vacancy content measured by XPS. Finally, the $\text{Ce}_{0.95}\text{La}_{0.05}\text{O}_8$ monolithic catalyst also shows an excellent durability of more than 100 hours, which is mainly due to the doping effect of lanthanum into the ceria oxides tailoring the structure and surface properties of the catalyst.

Keywords: La-doped CeO_2 Nanoparticles, Dimethyl Carbonate, Monolithic Catalysts, High Durability, CO_2 Utilization

1. Introduction

With the development of industry and extensive use of petrochemical resources, the increased greenhouse gas emissions have caused shocking climate change, which causes huge environmental problems [1], such as global warming. Carbon dioxide (CO_2), one of the most important greenhouse gases causing global warming, is a potential carbon resource as well [2]. Therefore, utilizing greenhouse gas CO_2 as an effective carbon resource for fixation and resource utilization is of great significance in alleviating the crisis of carbon sources and environmental protection [1].

Dimethyl carbonate (DMC), as a chemical raw material with low toxicity, excellent environmental protection performance as well as wide application, has broad potential application

prospects [3]. It can replace traditional carbonylation and methylation reagent and avoid the toxicity of phosgene, methyl chloroformate and dimethyl sulfate [4-7]. In addition, as a gasoline additive, DMC can improve its octane number and oxygen content, thus improving its antiknock performance [4, 6-8]. Based on the wide application and large demand of DMC, researchers recently have discovered a series of methods for the DMC synthesis, including carbonylation of methanol [9, 10], transesterification of carbonates [11, 12], alcoholysis of urea [13] and direct synthesis from CO_2 and methanol [14-17]. Among all of these routes, $\text{CO}_2 + 2\text{CH}_3\text{OH} \rightarrow \text{CH}_3\text{OC}(=\text{O})\text{OCH}_3 + \text{H}_2\text{O}$, this synthesis method has gradually become the most attractive method, which can realize the efficient use of the greenhouse gas CO_2 [17]. However, this direct synthetic method still faces some problems to be solved

urgently. The productivity of dimethyl carbonate is relatively low (e.g., <5%) due to the kinetic inertness and high thermodynamic stability of CO₂, as well as the by-product water can easily inactivate the catalyst [16], which leads to the poor reaction stability during the DMC synthesis. Therefore, we consider to overcome these problems by developing novel catalysts to effectively remove the generated water in the reaction system to improve the stability of the catalysts.

It has been reported that many nano-catalysts were effective for increasing the yield of this reaction, such as supported Cu-Ni bimetallic catalysts [18-20], organometallic compounds [21], zirconium oxide [22], cerium oxide [23, 24], CeO₂-ZrO₂ [5, 25, 26], and so on. According to the catalysts reported in the literature, transition metal-doped ceria catalysts showed excellent catalytic activity [5]. For example, Wang et al. [27] observed that the exposed crystal plane, defect sites and acid-base sites had a crucial impact on the well-defined CeO₂ nanocrystal catalyst. In addition, researchers also observed that doping transition metal elements (such as Ti [14, 16], Zn [28], Ca [29, 30], and La [23]) into CeO₂ will affect its storage/release oxygen capacity, defects stability as well as oxygen atom mobility, and thus to improve the catalytic activity and thermal stability of the catalyst.

Although the ceria-based catalysts have excellent catalytic performance for the DMC direct synthetic method compared with other catalysts, the CH₃OH conversion is still low (e.g., <5%), which seriously hinders its real application. This is mainly due to the water formed during the catalytic reaction can't be removed in time, so that the equilibrium of the reversible reaction shifts to the reverse reaction direction (according to Le Chatelier's principle) [14]. As a result, adding a dehydrating agent such as 2-cyanopyridine is of great necessity for high methanol conversion and DMC productivity [23]. However, dehydrating agents are generally expensive, highly toxic, and environmentally unfriendly. Therefore, to develop a low-cost, efficient and water-resistant catalyst to enhance the catalytic activity is worthwhile and urgently needed.

In this work, we prepare lanthanum-doped CeO₂ nanoparticles through a co-precipitation method. Nanocomposites of Ce_{1-x}La_xO₈ (where, x=0.00, 0.05, 0.10, 0.15, and 0.20) are coated on a honeycomb ceramic substrate to form Ce_{1-x}La_xO₈ monolithic catalysts for the direct synthesis of DMC. The monolithic catalysts have the virtue of lower pressure drop, larger BET surface area, better mechanical stability and thermal stability in compression to corresponding particulate catalysts [14, 31, 32]. In addition, the reaction products (water and DMC) produced by this reaction can be removed in time to improve the catalytic performance. The experimental results show that the lanthanum doped ceria monolithic catalysts exhibit excellent catalytic activity and great stability in DMC direct synthetic method in the absence of any dehydrating agents.

2. Experimental

2.1. Materials

All the reactants were analytical purity and used without any

further treatment except methanol was Chromatographically purity. Cerium ammonium nitrate ((NH₄)₂Ce(NO₃)₆·6H₂O) and Lanthanum nitrate (La(NO₃)₃·6H₂O) were purchased from Aladdin Co., Ltd. (Shanghai, China). Ethanol and urea were purchased from Chengdu Kelong Chemical Reagent Factory (Chengdu, China), Methanol (≥99.9%) was purchased from Shanghai Titan Technology Co., Ltd. (Shanghai, China). CO₂ (≥99.9%) and N₂ (≥99.9%) were obtained from Chengdu Xindu District Zhengrong Gas Co., Ltd. (Chengdu, China).

2.2. Preparation of Ce_{1-x}La_xO₈ Monolith Catalysts

Ce_{1-x}La_xO₈ nanoparticles were synthesized through co-precipitation method. The specific method for synthesis of Ce_{0.8}La_{0.2}O₈ nanoparticles was as follows: 15.00g of (NH₄)₂Ce(NO₃)₆·6H₂O and 2.96g of La(NO₃)₃·6H₂O were dissolved in deionized water to form a solution, which was mixed with the preformed urea aqueous solution prepared by 70g urea and transferred the mixture to a 1000 mL three-necked flask. The mixture was naturally heated in water bath to 90 °C for 5 hours under mechanical stirring. After the reaction, it was cooled, filtered, and washed with pure water and ethanol for several times. The collected precipitate was dried at 80 °C for more than 12 h and subsequently calcined at 400 °C for 4 h to obtain Ce_{0.8}La_{0.2}O₈ nanoparticles. Ce_{1-x}La_xO₈ (0 ≤ x ≤ 0.20) nanoparticles with different La dopants were prepared by changing the amounts of (NH₄)₂Ce(NO₃)₆·6H₂O and La(NO₃)₃·6H₂O in the starting solution.

Firstly, the as-obtained Ce_{1-x}La_xO₈ nanoparticles powder was milled with appropriate amount of pseudo-boehmite, deionized water and glacial acetic acid to prepare a slurry. Then, the cordierite honeycomb ceramic substrate (Figure 1, left), which has been manually polished into a specific shape, was immersed in the slurry, and the excess slurry was blown off with compressed air. Ultimately, the coated catalyst was dried at 80°C overnight and calcined at 400°C for 4 hours to obtain a Ce_{1-x}La_xO₈ monolith catalyst (Figure 1, right). The dried coating quality of Ce_{1-x}La_xO₈ had a mass of about 0.5 g.

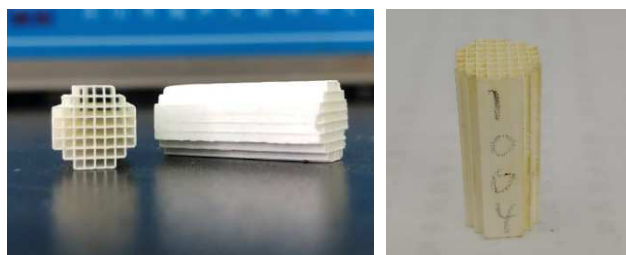


Figure 1. The appearance picture of cordierite honeycomb ceramic substrate (left) and monolithic catalyst (right, cordierite honeycomb ceramic substrate coated with Ce_{0.95}La_{0.05}O₈ nanoparticle). The pore density of cordierite honeycomb ceramic matrix ≥ 400cps.

2.3. Characterization Techniques

Transmission electron microscopy (TEM) was used to observe the micro morphology of the catalyst and measure the particle size. The TEM characterization was performed on a Philips FEI Tecnai G2 Sprit F30S-Twin transmission electron microscope with at an acceleration voltage of 120 kV. Sample

preparation process: firstly, the sample was ground into powder with an agate mortar, and then the ground sample was evenly dispersed in absolute ethanol, and then the suspension was dropped on the carbon film of copper mesh and dried in the air for testing. The particle size of these samples was determined by a random selection of >400 particles in different area of the TEM images. X-ray diffraction (XRD) was used to identify and analyze the phase, cell parameters and crystal form of the catalytic materials. Firstly, the samples were ground and pressed, and then put into a X Pert PRO MPD diffract meter (PANalytical B. V.) for X-ray diffraction analysis with a Cu K α ray ($\lambda=0.15418\text{nm}$) operated at 40 kV and 200 mA. The scanning range was $20^\circ \sim 80^\circ$ and the intensity was recorded by scintillation counter. The crystal phases were identified based on JCPDS card and Jade-6.5 X ray data processing software was used for data processing.

N₂ adsorption-desorption isotherms were measured on a Quanta chrome autos orb iQ instrument. Nitrogen was used as the adsorbed gas, and the analysis temperature was liquid nitrogen temperature. The sample was purged at 90 °C for 1 h, then vacuum pretreated at 300 °C for 3 h, and then cooled to room temperature for BET test. The specific surface area is calculated by the BET method, and the pore structure distribution map is obtained by the BJH method.

Raman spectra were recorded by a Renishaw in Via Raman

spectrometer. The scanning wavelength range is 100-1000 cm⁻¹. The light source is visible light source, and the wavelength is 532 nm.

The temperature-programmed reduction by hydrogen (H₂-TPR) of the catalyst was carried out in an AutochemII2920 instrument (Micromeritics) with a conventional TCD detector. Firstly, 100 mg of the sample was placed in a U-shaped quartz tube, and the sample was cleaned with 30 ml min argon at 200°C for 30 min. After cooling to 30°C under an argon atmosphere (30 mL/min), 10% H₂/Ar (30 mL/min) was introduced to the system and the temperature was programmed at a rate of 10°C/min to 900°C. The signal was recorded with a thermal conductivity detector, and finally dropped to room temperature.

The XPS characterization was analyzed by a Theta Probe system (Thermo Scientific) photoelectron spectroscopy using Al K α radiation at 13 KV and 20 mA. The electro binding energy was corrected by the C 1s (284.8 eV) contaminated carbon. To analyze the individual contributions of the Ce 3d and O 1s core levels, the absorption peaks of Ce 3d and O 1s were fitted. According to equation (1) and (2), the concentration of Ce³⁺ ($C[\text{Ce}^{3+}]$ %) and surface oxygen vacancies ($C[\text{O}_v]/\%$) of Ce_{1-x}La_xO₈ catalysts were calculated by semi-quantitative analysis of the integral peak area of respective valence state.

$$C[\text{Ce}^{3+}]\% = \frac{S_{\text{Ce}^{3+}}}{S_{\text{Ce}^{3+}} + S_{\text{Ce}^{4+}}} \times 100\%, \quad (1)$$

$$S_{\text{Ce}^{3+}} = S_{v'} + S_{u'}, \quad S_{\text{Ce}^{4+}} = S_v + S_{v''} + S_{v'''} + S_u + S_{u''} + S_{u'''}$$

$$C[\text{O}_v + \text{O}_c]\% = \frac{S_{\text{O}_v} + S_{\text{O}_c}}{S_{\text{O}_v} + S_{\text{O}_c} + S_{\text{O}_L}}, \quad (2)$$

$S_{\text{Ce}^{3+}}$: Sum of peak areas of Ce³⁺ on catalyst surface;

$S_{\text{Ce}^{4+}}$: Sum of peak areas of Ce⁴⁺ on catalyst surface

S_{OL} : Lattice oxygen area of catalyst surface.

S_{Oc} : Area of chemically adsorbed oxygen on catalyst surface.

S_{Ov} : Oxygen vacancy area on catalyst surface.[1].

2.4. Catalytic Performance Test

Direct synthesis DMC from CO₂ and CH₃OH was carried out in a continuous fixed-bed reactor. The prepared monolithic Ce_{1-x}La_xO₈ catalyst was placed into a stain reactor tube and sealed, and then purged with CO₂ flow for more than 10 min to drain the internal air. After the reaction system was heated to the required temperature, a constant current pump was used to

introduce CH₃OH into the system to participate in the reaction. During the experiment, the flow rate of CO₂ (40 mL·min⁻¹) and CH₃OH (0.145 mL·min⁻¹, it is calculated from the molar ratio of CH₃OH: CO₂=2:1 and the flow of CO₂) remained constant, which can be precisely controlled by evaporating CH₃OH and CO₂ flow. The tail gas component after the reaction was analyzed online using a gas chromatography instrument (Agilent GC 7890B) equipped with a hydrogen flame ionization detector (FID) for the detections of organic species (e.g. CH₃OH, DMC, HCHO, DME, et al.) and a thermal conductivity detector (TCD) for CO and CO₂ [2, 3]. The calculation formula for CH₃OH conversion (conversion) and DMC selectivity is:

$$\text{CH}_3\text{OH conv.}(\%) = \frac{3C_{\text{DMC}} + 2C_{\text{DME}} + C_{\text{HCHO}} + C_{\text{CO}}}{C_{\text{CH}_3\text{OH}} + 3C_{\text{DMC}} + 2C_{\text{DME}} + C_{\text{HCHO}} + C_{\text{CO}}} \times 100\% \quad (3)$$

$$\text{DMC sel.}(\%) = \frac{3C_{\text{DMC}}}{3C_{\text{DMC}} + 2C_{\text{DME}} + C_{\text{HCHO}} + C_{\text{CO}}} \times 100\% \quad (4)$$

where, C_i represents the concentration of a component (i).

3. Results and Discussion

3.1. Characterization of Ce_{1-x}La_xO₈ Composites

The microstructure and the size distribution of the

Ce_{1-x}La_xO₈ ($x=0.00, 0.05, 0.10, 0.15$, and 0.20) nanoparticles are depicted in Figure 2, respectively. It is noted that the diameter of Ce_{1-x}La_xO₈ nanoparticles varies from 3~7 nm with the molar ratios of La/Ce. In addition, we can also notice with the increasing concentration of lanthanum dopant, the size of the Ce_{1-x}La_xO₈ nanoparticles slightly decrease from 6.8 nm (CeO₂) to 3.8 nm (Ce_{0.80}La_{0.20}O₈). This phenomenon is in good line with XRD results (Figure 3a and Table 1).

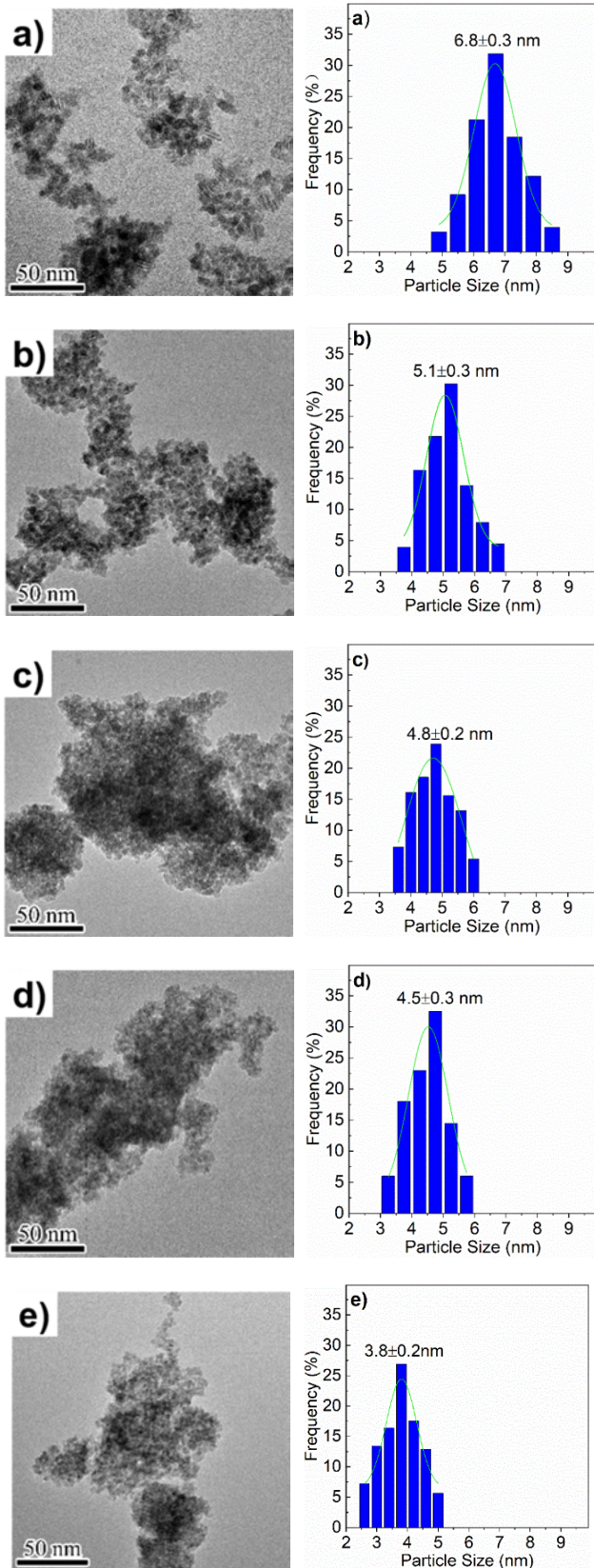


Figure 2. TEM and particle size distribution images of $Ce_{1-x}La_xO_8$

nanoparticles.

Figure 3a shows the XRD patterns of the $Ce_{1-x}La_xO_8$ nanoparticles. Obviously, lanthanum doped ceria composites showed four diffraction peaks in the range of $2\theta=20 \sim 60$, where $2\theta=28.5^\circ$, 33.0° , 47.4° , and 56.2° , which were just corresponding to (111), (200), (220) and (311) crystal planes of CeO_2 . [14]. Further, no peaks are observed corresponding to individual lanthanum oxide, which can be explained by the formation of $Ce_{1-x}La_xO_8$ solid solution. Figure 3b shows that the $Ce_{1-x}La_xO_8$ nanoparticles exhibit broader (111) diffraction peak which shift to lower 2θ with the increasing concentration of lanthanum species. Moreover, we also notice the (111) lattice parameter of $Ce_{1-x}La_xO_8$ nanoparticles increases from 0.5413 nm (CeO_2) to 0.5453 nm ($Ce_{0.80}La_{0.20}O_8$, Table 1) as a result of La^{3+} ion radius (0.106 nm) is larger than Ce^{4+} ion radius (0.097 nm). When La^{3+} enters in the ceria lattice to replace partial Ce^{4+} , the lattice parameter of ceria will increase, and these changes are consistent with the previous reports [14]. All the results indicate that La^{3+} is successfully introduced into Ceria lattice, forming a solid solution and increasing the interplanar crystal spacing [33], which increases from 0.3127 nm for CeO_2 nanoparticles to 0.3148 nm for $Ce_{0.80}La_{0.20}O_8$ nanoparticles.

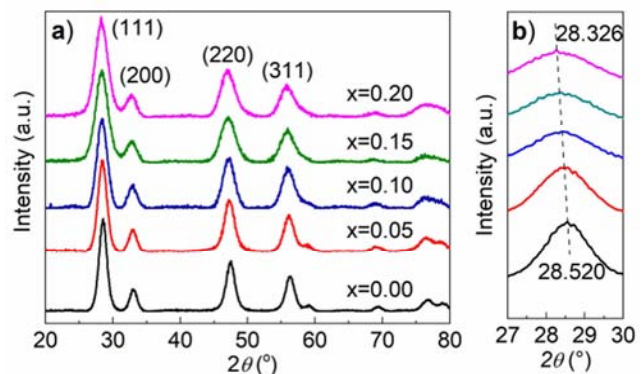


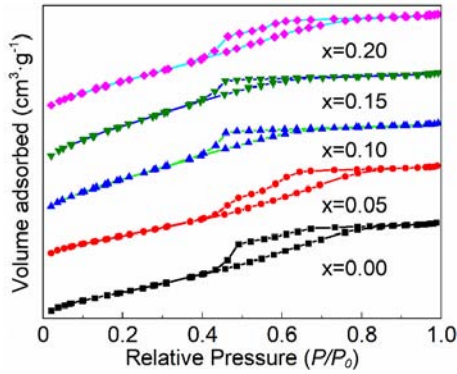
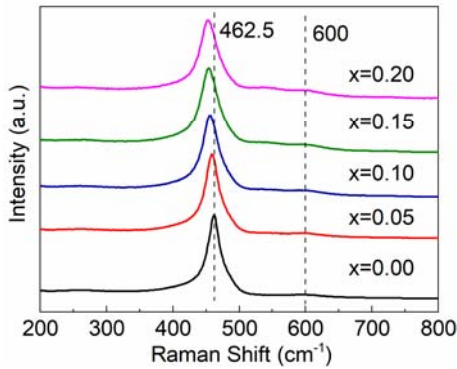
Figure 3. (a) XRD patterns of the $Ce_{1-x}La_xO_8$ composites. (b) Zoom-in of the (111) peak.

The adsorption equilibrium isotherms are also measured to analyze the pore volume, average pore size and specific surface area of the catalyst. From the adsorption equilibrium isotherms of the prepared $Ce_{1-x}La_xO_8$ composite oxides, we can see all of them show type IV adsorption equilibrium isotherms and exhibit obvious H_2 hysteresis loop in the IUPAC classification begins to appear with the relative pressure (P/P_0) 0.4 as the apex, indicating that they are typical mesoporous materials (Figure 4). The pore volume, average pore size and Brunauer–Emmett–Teller (BET) surface area of the synthesized $Ce_{1-x}La_xO_8$ composites are summarized in Table 1. Obviously, the BET surface area of $Ce_{1-x}La_xO_8$ samples increase when more lanthanum ions are induced into the ceria and $Ce_{0.95}La_{0.05}O_8$ shows the largest pore volume and average pore size.

Table 1. Structural and textural properties of $Ce_{1-x}La_xO_8$ nanoparticles.

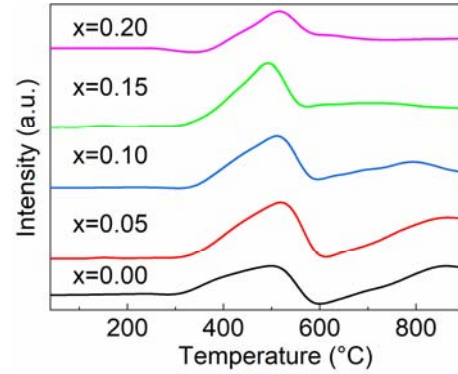
samples	(111) plane		Lattice (nm)	Crystalline size ^b	S_{BET} ($m^2 \cdot g^{-1}$)	V_{Pore} ($cm^3 \cdot g^{-1}$)	Average pore size (nm)
	2 ($^\circ$)	Parameter ^a					
CeO_2	28.520	0.3127	0.5413	6.8	142.9	0.141	3.937
$Ce_{0.95}La_{0.05}O_8$	28.456	0.3134	0.5429	5.1	159.0	0.163	4.104
$Ce_{0.90}La_{0.10}O_8$	28.388	0.3141	0.5437	4.8	158.6	0.119	3.003
$Ce_{0.85}La_{0.15}O_8$	28.361	0.3144	0.5446	4.5	171.8	0.122	2.847
$Ce_{0.80}La_{0.20}O_8$	28.326	0.3148	0.5453	3.8	186.0	0.160	3.437

^a The lattice parameter is calculated using Vegard's law. ^b The size of the oxide nanoparticles is estimated by TEM.

**Figure 4.** N_2 adsorption-desorption isotherm of $Ce_{1-x}La_xO_8$ nanocomposites.**Figure 5.** Raman spectra of $Ce_{1-x}La_xO_8$ nanoparticles

Further, the $Ce_{1-x}La_xO_8$ nanoparticles are characterized by Raman spectroscopy, as seen in Figure 5. The sharp peak at 462.5 cm^{-1} is the F_{2g} vibration peak of Ce–O [34], indicating that $Ce_{1-x}La_xO_8$ nanoparticles solid solution are cubic fluorite structure, matching well with the above XRD results. According to previous studies [35, 36], oxygen vacancies should be the catalytically active mediators during the DMC synthesis. With the increasing La^{3+} content, the peak intensity associated with the defect around 600 cm^{-1} increases slightly, indicating that the doping of La^{3+} promotes the formation of more defects and conforming the potential of the $Ce_{1-x}La_xO_8$ composites for DMC production. It is further basically consistent with the XPS results. Moreover, the F_{2g} bands become asymmetric, broaden with the increasing content of La^{3+} increasing and gradually shift to low wavenumber compared to pure ceria, which mainly because La^{3+} replacing part of Ce^{4+} will increase the bond length of Ce–O, causing the Raman spectrum moving to lower energy [5, 37, 38]. According to the literature, we have learned that this red shift is concerned with the change of the force between atoms,

which is susceptible to the lattice spacing and the change of bond length [5].

**Figure 6.** H_2 -TPR profiles of $Ce_{1-x}La_xO_8$ nanoparticles.

Next, Figure 6 shows the H_2 -TPR curve of the $Ce_{1-x}La_xO_8$ nanoparticles. CeO_2 nanoparticles show two obvious reduction peaks at $\sim 504^\circ\text{C}$ and $\sim 864^\circ\text{C}$, respectively. The reduction peak at 504°C is related to the surface reduction, while the peak at higher reduction temperature is ascribed to bulk reduction [39, 40]. And the surface hydrogen consumption of CeO_2 nanoparticles is $0.666\text{ mmol} \cdot g^{-1}$. Under the condition of hydrogen reduction, oxygen species on the surface of coordination unsaturated can be easily removed at a lower temperature. However, the bulk oxygen anions can only be reduced at a relative higher temperature after being transported to the surface [33]. When La^{3+} is doped into ceria, the corresponding hydrogen consumption under low temperature conditions first increases and then slightly decreases. It shows the doping of an appropriate amount of La_2O_3 can increase the amount of surface oxygen species that can be removed. Moreover, we also find that La^{3+} doping can cause the bulk reduction peak of CeO_2 move to the low temperature [41]. For $Ce_{0.95}La_{0.05}O_8$, $Ce_{0.90}La_{0.10}O_8$ and $Ce_{0.85}La_{0.15}O_8$ nanoparticles, the bulk reduction peaks decrease to 845°C , 794°C and 703°C respectively. For $Ce_{0.80}La_{0.20}O_8$ nanoparticles, at about 513°C , the bulk reduction peak and the surface reduction peak almost completely merge into a wide peak, which is consistent with the reported results in literature that the high temperature and low temperature reduction peaks of $Ce_{1-x}La_xO_8$ composites gradually merged into one wide peak with the increase of lanthanum content [33]. Furthermore, CeO_2 and $Ce_{1-x}La_xO_8$ nanoparticles have the same the bulk reduction mechanism. However, the bulk reduction peak of $Ce_{1-x}La_xO_8$ nanoparticles shifts to low temperature may be due to the relatively rapid

diffusion of oxygen from bulk to the surface, the bulk oxide ions continue to move to the surface, forming more active hydrogen adsorption sites on the surface [41].

The chemical state of Ce on the surface of $\text{Ce}_{1-x}\text{La}_x\text{O}_\delta$ nanoparticles is studied by XPS. As shown in Figure 7a, the content of La has an obvious effect on the chemical state of Ce on the catalyst surface. According to the literature method, we can decompose the Ce 3d spectrum into eight peaks [5]: u''' , u'' , u' , u , v''' , v'' , v' , and v . The four v bands belong to the spin-orbit splitting peaks of Ce 3d_{5/2}, while the other four u bands are attributed to the spin-orbit splitting peaks of Ce 3d_{3/2}. The six peaks of u , u'' , u''' , v , v'' and v''' are the final state spin orbit splitting peaks of Ce^{4+} , and the two peaks of u' and v' are the final state spin orbit splitting peaks of Ce^{3+} , which indicates that there are two valence states of Ce on the surface of $\text{Ce}_{1-x}\text{La}_x\text{O}_\delta$ nanoparticles. Quantitative analysis of Ce^{3+} on the catalyst surface is carried out and the calculation method is shown in 2.3 in Chapter 2. It is showed that the Ce^{3+} concentration changes with the change of lanthanum doping and $\text{Ce}_{0.95}\text{La}_{0.05}\text{O}_\delta$ nanoparticles exhibit the highest Ce^{3+} concentration of 21.13% among all the subjects. Moreover, $\text{Ce}_{1-x}\text{La}_x\text{O}_\delta$ nanoparticles exhibit higher $\text{Ce}^{3+}/\text{Ce}_{\text{total}}$ proportion compared with pure ceria, which indicates that the doping of La significantly increases the surface Ce^{3+} concentration. Zhang et al [42] has found the generation of Ce^{3+} leads to the formation of oxygen vacancies. As Ce^{3+} concentration increases, oxygen vacancies concentration increases as well [5]. As a result, $\text{Ce}_{0.95}\text{La}_{0.05}\text{O}_\delta$ nanoparticles may have the highest surface oxygen vacancy concentration.

The O 1s spectra of the prepared composites is shown in Figure 7b. According to the literature method, the XPS spectrum of O 1s can be divided into 3 peaks after fitting: O_L (~529.3 eV), O_V (~530.5 eV), and O_C (~532.2 eV), which correspond to lattice oxygen species, surface oxygen vacancies and chemisorbed oxygen species, respectively [14, 43]. It is found that the concentration of Ce^{3+} and O_V varied with the doping amount of lanthanum. Table 2 indicates O_V levels are 26.75%, 29.35%, 28.83%, 28.51% and 27.57%,

corresponding to CeO_2 , $\text{Ce}_{0.95}\text{La}_{0.05}\text{O}_\delta$, $\text{Ce}_{0.90}\text{La}_{0.10}\text{O}_\delta$, $\text{Ce}_{0.85}\text{La}_{0.15}\text{O}_\delta$ and $\text{Ce}_{0.80}\text{La}_{0.20}\text{O}_\delta$, respectively, which indicate that doping appropriate amount of lanthanum contributes to the increase of Ce^{3+} content on the catalyst surface, which is mainly because the chemical state of La^{3+} is different from that of Ce^{3+} . After doping lanthanum, in order to maintain the overall electronegativity of CeO_2 , more trivalent cerium is produced, which leads to the increase of oxygen vacancies, which is consistent with the previous Raman results. In addition, the survey spectra of $\text{Ce}_{1-x}\text{La}_x\text{O}_\delta$ composites are shown in Figure 8, there is no impurity peak except carbon, because the peak arrangement in the XPS spectra is performed using the characteristic peak of C 1s as a reference (binding energy is 284.8 eV).

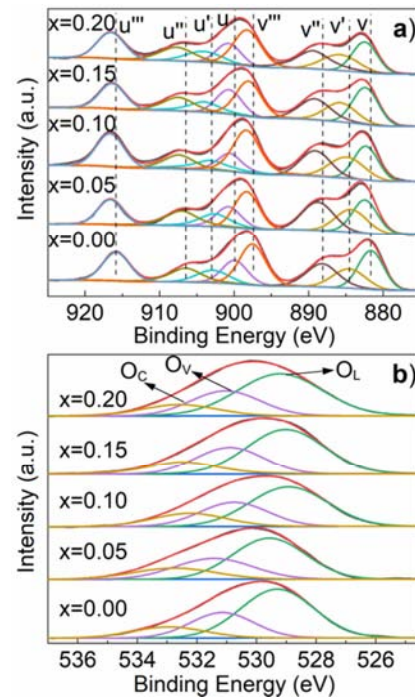


Figure 7. XPS spectra of the $\text{Ce}_{1-x}\text{La}_x\text{O}_\delta$ nanocomposites. (a) Ce 3d; (b) O 1s.

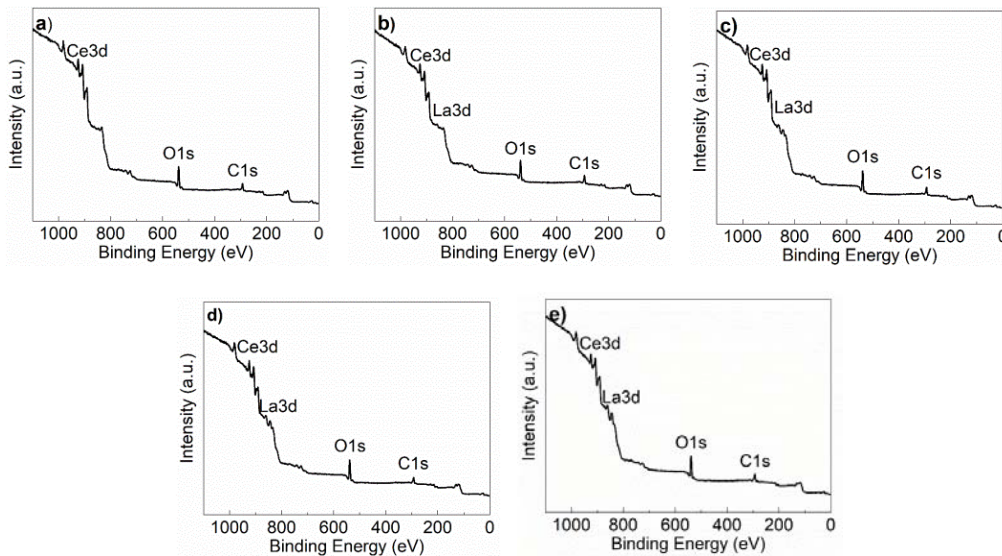


Figure 8. XPS survey of $\text{Ce}_{1-x}\text{La}_x\text{O}_\delta$ nanoparticles: (a) $x=0.00$, (b) $x=0.05$, (c) $x=0.10$, (d) $x=0.15$, (e) $x=0.20$.

Table 2. Relative Ratios of the Surface Concentration of Ce^{3+} and Oxygen Vacancy Based on XPS Analysis.

Samples	Concentration of Ce^{3+} (%)	Concentration of O_v (%)
CeO_2	17.54	26.75
$Ce_{0.95}La_{0.05}O_8$	21.13	29.35
$Ce_{0.90}La_{0.10}O_8$	20.34	28.83
$Ce_{0.85}La_{0.15}O_8$	19.32	28.51
$Ce_{0.80}La_{0.20}O_8$	18.18	27.57

3.2. Catalytic Performance

The activity evaluation results of the $Ce_{1-x}La_xO_8$ monolithic catalyst for catalyzing CO_2 and methanol synthesis DMC are shown in Figure 9. It is carried out on a continuous tubular fixed-bed reactor. Firstly, $Ce_{0.95}La_{0.05}O_8$ nanoparticles are chosen as the catalyst to synthesize DMC directly by introducing CH_3OH and CO_2 into the continuous tubular fixed-bed reactor. The catalytic performance of $Ce_{0.95}La_{0.05}O_8$ nanoparticles coated on the honeycomb ceramics ($Ce_{0.95}La_{0.05}O_8$ monolithic catalyst) is investigated under the condition of reaction temperatures of $100^\circ C \sim 180^\circ C$ and reaction pressure of 2.4 MPa without any dehydrating agent. It is observed from Figure 9a that when the reaction temperature increases from $100^\circ C$ to $180^\circ C$, the conversion of CH_3OH increases firstly and then decreases slightly and it reaches the maximum of 22.01% when the reaction temperature is at $140^\circ C$. Besides, as the temperature increases, by-products increase as well, and the selectivity of dimethyl carbonate gradually decreases, which mainly due to the observation of two competing and side reactions during the catalytic process: (1) $2CH_3OH \rightarrow CH_3OCH_3 + H_2O$ and (2) $CH_3OH + CO_2 \rightarrow HCHO + CO + H_2O$. As the reaction temperature increases, the above two side reactions are promoted, resulting in more by-products (DME, HCHO and CO). [44-46, 19].

Furthermore, we compare the catalytic performance of the $Ce_{0.95}La_{0.05}O_8$ monolithic catalyst with the $Ce_{0.95}La_{0.05}O_8$ nanoparticle catalyst which are simply filled in a continuous tubular fix-bed reactor ($Ce_{0.95}La_{0.05}O_8$ granular catalyst, Figure 9b) as well. The $Ce_{0.95}La_{0.05}O_8$, coated on the honeycomb ceramics, gives a higher CH_3OH conversion (Figure 9a, black lines) and DMC selectivity (Figure 9a, red lines) than that of $Ce_{0.95}La_{0.05}O_8$ granular catalyst at each reaction temperatures (e.g., $100-180^\circ C$) in the direct dimethyl carbonate synthesis from carbon dioxide, indicating that the $Ce_{0.95}La_{0.05}O_8$ monolithic catalyst give much better catalytic performance than the corresponding granular catalyst. This is mainly due to the following facts that the by-product water can be quickly detached from the surface of

the monolithic catalyst during the reaction process, thus avoiding the catalyst activity reduction or even deactivation due to water poisoning. In all, the catalysts uniformly coated onto the honeycomb ceramics indeed can largely improve the catalytic performance (including the CH_3OH conversion and selectivity toward target product dimethyl carbonate) in the direct dimethyl carbonate synthesis reactions investigated in the direct DMC synthesis at $140^\circ C$, Figure 10a. The catalytic activity of the $Ce_{1-x}La_xO_8$ catalysts is $Ce_{0.95}La_{0.05}O_8 > Ce_{0.90}La_{0.10}O_8 > Ce_{0.85}La_{0.15}O_8 > Ce_{0.80}La_{0.20}O_8 > CeO_2$ with the DMC yield of 18.15%, 15.81%, 15.01, 13.31 and 12.34%. Interestingly, a volcano-type curve of the catalytic activity is observed, the $Ce_{0.95}La_{0.05}O_8$ nanocomposite catalyst shows the best catalytic activity with the highest CH_3OH conversion of 22.01% and DMC selectivity of 82.53%. Moreover, the activity of these $Ce_{1-x}La_xO_8$ nanocomposites is closely related to the concentration of the surface $[O_C + O_v]$ species, consistent with the previous observation [14]. It is worth to note that the DMC selectivity is gradually decreases from 85.86% for CeO_2 to 78.24% for $Ce_{0.80}La_{0.20}O_8$ with the increasing lanthanum dopants. In general, $Ce_{0.95}La_{0.05}O_8$ monolithic catalyst exhibits the highest CH_3OH conversion and DMC productivity among all the lanthanum doped cerium oxide. Furthermore, Table 3 compares $Ce_{0.95}La_{0.05}O_8$ monolithic catalyst with some previously reported granular catalysts. We note that $Ce_{0.95}La_{0.05}O_8$ monolithic catalyst exhibits higher catalytic activity than that of cerium-based particulate catalysts without dehydrating agents and other commonly used catalyst in the literatures [8, 16, 20, 47, 48]. This is mainly attributed to the monolithic catalysts can reduce the interaction between catalyst and gas (including reactants and products) and remove the generated water in time, thus improving the catalytic activity.

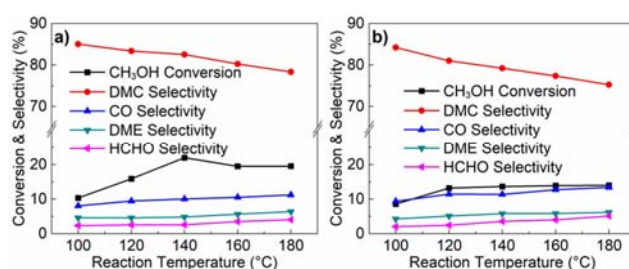


Figure 9. Comparison of the catalytic performance of the $Ce_{0.95}La_{0.05}O_8$ catalyst ((a) coated on the honeycomb ceramics and (b) simply filled with 40-60 mesh) Reaction conditions: amount of catalyst=500 mg, $CH_3OH:CO_2=2.0:1.0$ (molar ratio), $P=2.4$ MPa, $T=140^\circ C$ [14].

Table 3. Summary of reaction conditions and catalyst activity over various catalyst using fixed-bed reactor without dehydrating agents.

Catalyst	T ($^\circ C$)	P (MPa)	CH_3OH Con (%)	DMC Sel. (%)	Reference
$Ti_{0.04}Ce_{0.96}O_2$	140	1.0	5.4	83.1	[16]
$H_3PW_{12}O_{40}/Ce_{0.1}Ti_{0.9}O_2$	170	-	5.5	91.4	[47]
$Zr_{0.10}Ce_{0.90}O_2$	140	7.5	11.2	9.6	[48]
$Cu-CeO_2$	170	5.0	0.9	2.0	[49]
$Cu-Ni/molecular\ sieve$	120	1.1	5.8	86.0	[8]
$Cu_xNi_y@POP-PPh_3$	160	2.4	10.5	80.0	[20]
$CuNi/ZIF-8$	110	2.0	12.8	50.0	[19]
$Ce_{0.95}La_{0.05}O_8$	140	2.4	22.0	82.5	This work

Durability is one of the most significant parameters to evaluate the catalytic performance of a catalyst. Therefore, we have tested the durability of the $\text{Ce}_{0.95}\text{La}_{0.05}\text{O}_\delta$ monolithic catalyst prepared in this experiment as well. It is continuously examined more than 115 hours under 140°C and 2.4 MPa and the test result is noted in Figure 10b. It shows the CH_3OH conversion and DMC selectivity are almost constant at $\sim 22\%$ and $\sim 80\%$, respectively, and no deactivation is observed during the catalytic test of the initial 100 hours, which indicates that the durability of the $\text{Ce}_{0.95}\text{La}_{0.05}\text{O}_\delta$ monolithic catalyst is as high as 100 hours, which is much higher than that of $\text{Cu}_1\text{Ni}_1@\text{POP-PPH}_3$ monolithic catalyst (about 16 hours) [20] and $\text{Ti}_{0.1}\text{Ce}_{0.9}\text{O}_2$ monolithic catalyst (less than 50 hours) [14]. However, after the 100 hours' catalytic reaction, the conversion of CH_3OH and the selectivity of DMC decrease slightly to 20% and 76%, respectively. This phenomenon is a result of the formation of carbonates covering the catalyst surface [14] in the long-time reaction. Therefore, we can find that the doping of lanthanum can significantly improve the stability and durability of the cerium-based composite oxide catalyst in the direct dimethyl carbonate synthesis process.

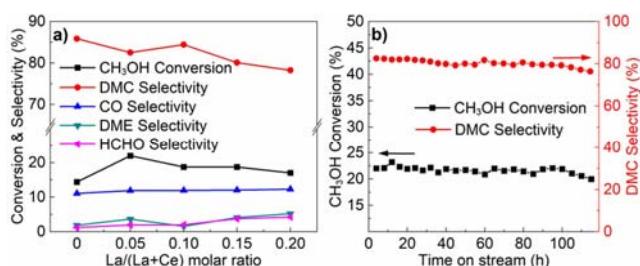


Figure 10. (a) Catalytic performance over the $\text{Ce}_{1-x}\text{La}_x\text{O}_\delta$ monolithic catalysts in DMC direct synthetic reaction. (b) Durability test of $\text{Ce}_{0.95}\text{La}_{0.05}\text{O}_\delta$ monolithic catalyst for a long time of over 115 hours. Reaction conditions: amount of catalyst=500 mg, $\text{CH}_3\text{OH}:\text{CO}_2=2.0:1.0$ (molar ratio), $P=2.4$ MPa, $T=140^\circ\text{C}$ [14].

4. Conclusions and Recommendation

In general, we have designed and prepared five kinds of La-doped spherical ceria nanoparticles via a co-precipitation method. H_2 -TPR, XPS and Raman spectra tests show that the introduction of lanthanum species has a certain effect on the surface oxygen vacancy sites of these $\text{Ce}_{1-x}\text{La}_x\text{O}_\delta$ nanocomposites. Besides, higher oxygen vacancy sites concentration is beneficial to the adsorption of CO_2 and improves the catalytic performance of DMC direct synthetic method. In addition, we also found that an appropriate amount of lanthanum doping into CeO_2 can significantly improve the reaction stability and durability. The $\text{Ce}_{0.95}\text{La}_{0.05}\text{O}_\delta$ monolithic catalyst can maintain high activity for more than 100 hours without obviously deactivation under the optimal reaction temperature (140°C).

Based on the above results and conclusions, the following suggestions are put forward.

This paper focuses on the experimental study of the effect of reaction conditions on the catalytic activity, without considering the internal mechanism of the reaction. It is

recommended to consider the reaction mechanism of the reaction in the later research, and reveal the internal connection between the catalytic activity and the structure of the catalyst by studying the reaction mechanism.

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