

The Potential Impact of Biofield Energy Treatment on the Physical and Thermal Properties of Silver Oxide Powder

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To cite this article:

Mahendra Kumar Trivedi, Rama Mohan Tallapragada, Alice Branton, Dahryn Trivedi, Gopal Nayak, Omprakash Latiyal, Snehasis Jana. The Potential Impact of Biofield Energy Treatment on the Physical and Thermal Properties of Silver Oxide Powder. *International Journal of Biomedical Science and Engineering*. Vol. 3, No. 5, 2015, pp. 62-68. doi: 10.11648/j.ijbse.20150305.11

Abstract: Silver oxide has gained significant attention due to its antimicrobial activities. The purpose of this study was to evaluate the impact of biofield energy treatment on the physical and thermal properties of silver oxide (Ag₂O). The silver oxide powder was divided into two parts, one part was kept as control and another part was received Mr. Trivedi's biofield energy treatment. The control and treated samples were analyzed using X-ray diffraction (XRD), differential scanning calorimetry (DSC), thermogravimetric analysis (TGA), and Fourier transform infrared (FT-IR) spectroscopy. The XRD diffractogram showed that the crystallite size of treated sample was significantly altered on the planes (200), (311), and (220) by 100, 150 and -25% respectively, with respect to control. The DSC result exhibited that the thermal energy required to decompose the silver oxide to silver and oxygen was altered from -12.47 to 71.58% in treated samples as compared to the control. TGA showed that the onset temperature of thermal degradation was reduced from 335°C (control) to 322.4°C. In addition, the rate of weight loss in treated sample was increased by 4.14% as compared to the control. Besides, the FT-IR did not show any alteration in absorption wavenumber of treated sample as compared to the control. Hence, the XRD, DSC and TGA data revealed that the biofield energy treatment has a significant impact on the physical and thermal properties of silver oxide powder. Therefore, the biofield energy treatment might improve the dissolution rate in formulation and bioavailability of treated silver oxide as compared to control.

Keywords: Silver Oxide, Biofield Energy Treatment, X-Ray Diffraction, Differential Scanning Calorimetry, Thermogravimetric Analysis, Fourier Transform Infrared Spectroscopy

1. Introduction

Silver (Ag) is a naturally occurring ductile and malleable element. It has higher thermal and electrical conductivity among all metals [1]. In addition, metallic silver has been used in several medical application which includes surgical prosthesis and splints, and coinage [2]. The antibacterial properties of silver is known to world from ancient times [3]. Recently, silver compounds such as silver chloride, silver nitrate, and silver oxide also gained significant attention due to their antimicrobial properties [4]. Allahverdiyev *et al.* reported that the silver oxide has an antibacterial effect against drug-resistant bacteria and leishmania parasites [5]. Also, the presence of silver oxide in glassy system showed its

significant effect against *Escherichia coli* and *Salmonella* and *Pseudomonas* bacteria [6]. It is also reported that the efficacy of drug is highly depends on their physical and thermal properties [7, 8]. Thus, it is important to study a treatment strategy *i.e.* biofield energy treatment, which may modify the physical and thermal properties of the silver oxide. The National center for Complementary and alternative medicine (NCCAM) has recommended uses of alternative CAM therapies (*e.g.* healing therapy) in the healthcare sector [9]. Besides, a human has the capability to harness the energy from the environment/Universe and transmit it to any object around the Globe. The object(s) receive the energy and respond into a useful way that is called biofield energy, and this process is known as biofield energy treatment. Mr.

Trivedi's unique biofield energy treatment (The Trivedi Effect®) had been studied in various scientific fields such as material science [10-12], microbiology [13] and agricultural [14]. The biofield energy treatment has significantly altered the physical and thermal properties in cadmium [15] and magnesium [16]. Recently, our group reported that biofield treatment has altered the bond length of Ti-O in barium titanate [17]. Therefore, based on excellent outcomes with biofield energy treatment on metals and ceramics, this work was undertaken to evaluate the effect of this treatment on the physical and thermal properties of the silver oxide powder using X-ray diffraction (XRD), differential scanning calorimetry (DSC), thermogravimetric analysis (TGA), and Fourier transform infrared (FT-IR) spectroscopy.

2. Materials and Methods

Silver oxide powder was procured from Sigma Aldrich, USA. The procured powder was divided into two parts coded as control and treated. The control part was remained the same and the treated part was in sealed pack, handed over to Mr. Trivedi for biofield energy treatment under standard laboratory condition. Mr. Trivedi provided the treatment through his energy transmission process to the treated sample without touching the sample. After that, the control and treated samples were characterized by using XRD, DSC, TGA, and FT-IR techniques.

2.1. XRD Study

The XRD analysis of control and treated silver oxide was performed on Phillips, Holland PW 1710 X-ray diffractometer system. The x-ray of wavelength 1.54056 Å was used. From the XRD diffractogram, peak intensity counts, d value (Å), full width half maximum (FWHM) (θ°), relative intensity (%) were obtained. The crystal structure parameters such as lattice parameter, unit cell volume of control and treated were computed using PowderX software.

The crystallite size (D) was calculated by using Scherrer equation as following:

$$D = k\lambda / (b \cos \theta)$$

Here, $\lambda = 1.54056$ Å, b is full width half maximum (FWHM) of peaks and k is the equipment constant (=0.94).

The percentage change in crystallite size was calculated using following formula:

$$\% \text{ change in crystallite size} = [(D_t - D_c) / D_c] \times 100$$

Where, D_c and D_t are crystallite size of control and treated powder samples respectively.

2.2. Thermal Analysis

The thermal analysis of silver oxide powder was performed using DSC and TGA-DTA. For DSC analysis, Pyris-6 Perkin Elmer DSC at a heating rate of 10°C/min with nitrogen atmosphere, was used. The thermal decomposition temperature and thermal energy of the samples were obtained

from the DSC curve. The percentage change in thermal energy (ΔH) was calculated using following equation:

$$\begin{aligned} &\% \text{ change in thermal energy } (\Delta H) \\ &= \frac{[\Delta H_{\text{Treated}} - \Delta H_{\text{Control}}]}{\Delta H_{\text{Control}}} \times 100 \end{aligned}$$

Where, $\Delta H_{\text{Control}}$ and $\Delta H_{\text{Treated}}$ are the latent heat of fusion of control and treated samples, respectively. Besides, for TGA analysis, Mettler Toledo simultaneous TGA-DTA was used to analyze the thermal characteristics of silver oxide powder. The samples were heated from room temperature to 900°C with a heating rate of 10°C/min under nitrogen atmosphere.

2.3. FT-IR Spectroscopy

FT-IR spectra were recorded on Shimadzu's Fourier transform infrared spectrometer (Japan) with frequency range of 4000-500 cm^{-1} . The analysis was accomplished to evaluate the effect of biofield treatment on dipole moment, force constant and bond strength in chemical structure.

3. Results and Discussion

3.1. XRD Study

The XRD is a non-destructive and quantitative technique, which has been extensively used to study the crystal structure and its parameters of a compound. The XRD diffractogram of control and treated silver oxide samples is shown in Figure 1. It can be observed that the control sample showed the crystalline peaks at Bragg angle (2θ) 32.72°, 37.95°, 54.78°, 65.28° and 68.6° which were attributed to the crystalline planes (111), (200), (220), (311), and (222) according to Joint committee on powder diffraction standards (JCPDS) file no. 12-0793 [18]. The treated sample also exhibited the similar peaks at Bragg angle 32.72°, 37.97°, 54.78°, 65.31°, and 68.59°. These peak suggested the cubic crystal structure of the silver oxide in control and treated sample. Furthermore, the crystallite size was computed on each crystalline planes and presented in Table 1. Data showed that the crystallite size of treated silver oxide was altered on the crystalline plane (200), (220) and (311). The crystallite size was found as 72.9 nm and 98.4 nm on planes (200) and (311) respectively in control sample, while it was significantly increased to 145.9 nm and 245.9 nm respectively in treated sample. Furthermore, the crystallite size on plane (220) was decreased from 155.4 nm (control) to 116.6 nm in treated silver oxide sample after biofield treatment. It indicated that the crystallite size of treated sample was significantly altered by 100%, -25%, and 150% on crystalline planes (200), (220), and (311) as compared to the control (Figure 2). However, the crystallite size on planes (111) and (222) were showed same crystallite size in control and treated sample. In cubic crystals system the planes (111) and (222) are parallel to each other, thus both planes showed similar behavior. Subhan *et al.* reported that the crystallite size of silver oxide

nanocomposite was increased after annealing at high temperature [19]. In addition, it was also reported that the reduction in dislocation density in silver oxide leads to increase the crystallite size [20]. Thus, it is assumed that the biofield energy treatment probably altered the dislocation density on plane (200), (220) and (311) and that might be responsible for the alteration in the crystallite size along these planes. Furthermore, it is reported that the alteration in crystallite size affects the solubility of drugs in fluids [21]. Thus, it is assumed that the alteration in crystallite size in treated sample might influence the solubility of silver oxide

in topical formulation, rate of penetration and bioavailability across the skin. Moreover, the XRD data of the control and treated silver oxides were analyzed using PowderX software and results are presented in Table 2. The data showed that lattice parameter and unit cell volume were slightly reduced in treated sample as compared to control. The density data showed minimal increase (0.029%), whereas molecular weight data showed minimal decrease (0.031%) in treated sample as compared to the control. Thus, crystal structure parameters did not show any significant alteration in treated sample as compared to the control.

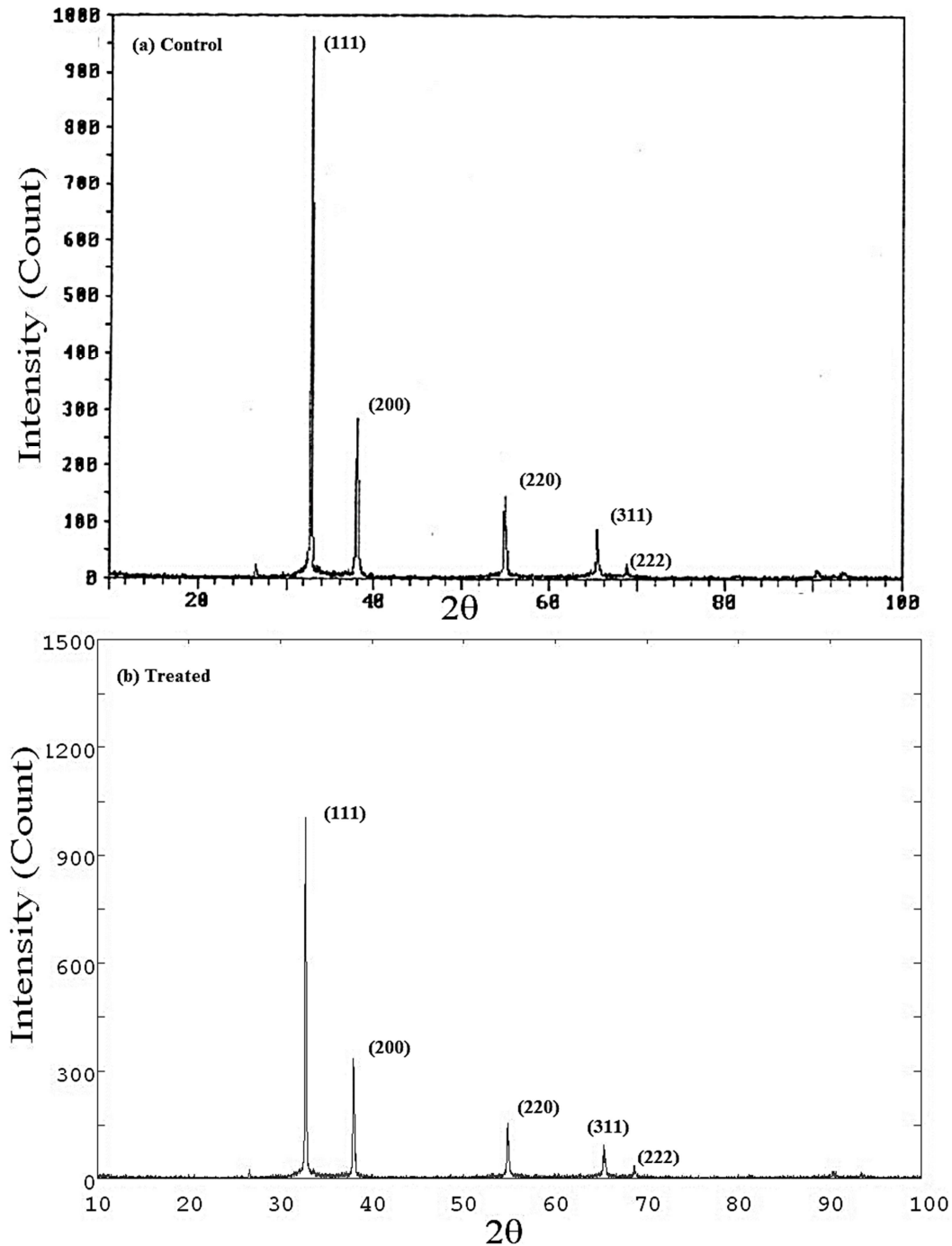


Fig. 1. X-ray diffractogram of silver oxide powder.

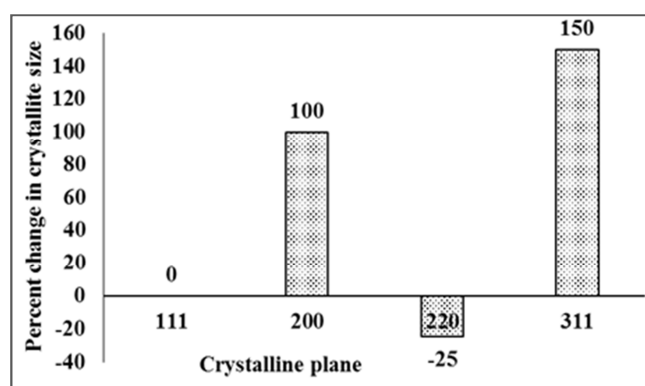
Table 1. The crystallite size and plane analysis of silver oxide powder.

Plane (hkl)	Control			Treated		
	Bragg's angle (2 θ)	FWHM (degree)	Crystallite size (nm)	Bragg's angle (2 θ)	FWHM (degree)	Crystallite size (nm)
111	32.72	0.12	86.3	32.72	0.12	86.3
200	37.95	0.14	72.9	37.97	0.08	145.9
220	54.78	0.08	155.4	54.78	0.10	116.6
311	65.28	0.12	98.4	65.31	0.06	245.9
222	68.60	0.10	125.3	68.59	0.1	125.3

FWHM: Full width half maximum of XRD peak

Table 2. Effect of biofield energy treatment on lattice parameter, unit cell volume density atomic weight, nuclear charge per unit volume of silver oxide powder.

Group	Lattice parameter (Å)	Unit cell volume ($\times 10^{-23}$ cm ³)	Density (g/cc)	Molecular weight (g/mol)
Control	4.7363	10.6250	7.31319	234.00145
Treated (T1)	4.7359	10.6217	7.31546	233.92877
Percent Change	-0.01	-0.030	0.029	-0.031

**Fig. 2.** Effect of biofield energy treatment on crystallite size along crystalline planes of silver oxide powder.

3.2. Thermal Analysis

Table 3. Differential scanning calorimetry (DSC) analysis of silver oxide powder.

Parameter	Control	T1	T2	T3
Onset (°C)	381.12	383.40	383.16	380.51
Peak (°C)	401.52	401.51	398.74	395.07
Endset (°C)	414.17	420.59	414.56	403.98
ΔH (J/g)	132.94	228.10	158.77	116.35
Percent change in ΔH	-	71.58	19.42	-12.47

The DSC analysis data of control and treated silver oxide samples are presented in Table 3. The DSC thermogram of control silver oxide showed a broad endothermic peak, which started from temperature (onset) 381.12 °C and ended at 414.17 °C. It is reported that the silver oxide decomposed to silver and oxygen at temperature around 400 °C [22]. Thus, the endothermic peak found in control sample was due to decomposition of silver oxide. The control sample showed the peak temperature for this process at 401.52°C. Furthermore, the treated samples T1, T2, and T3 were started to decompose at 383.40°C, 383.16°C, and 380.51°C respectively. It indicated that the onset temperature was slightly increased by 2.28°C and 2.06°C in T1 and T2 respectively, however it was slightly decreased by 0.61°C in

T3 as compared to the control. Moreover, the endset of this decomposition process was found at 420.59°C, 414.56°C, and 403.98°C in T1, T2, and T3 respectively. In addition, the peak temperature of this process was observed at 401.51°C, 398.74°C, and 395.07°C in treated samples T1, T2, and T3 respectively. The energy absorbed by the control sample in the decomposition of silver oxide was 132.94 J/g and it was changed to 228.10J/g, 158/77J/g, and 116.35J/g in treated in T1, T2, and T3, respectively. It indicated that the energy utilized in the decomposition of silver oxide in T1 and T2 samples was 71.58% and 19.42% higher as compared to the control. On the contrary, the energy utilized in the decomposition of T3 sample was 12.47% lower with respect to the control. Nevertheless, the increase in thermal decomposition energy in treated sample (T1 and T2) indicated that the thermal stability of silver oxide might increase after biofield energy treatment. Our group, previously reported that the biofield energy treatment had increased the latent heat of fusion in cadmium powder [15]. Thus, it is assumed that the energy, which probably transferred through biofield treatment, possibly altered the bonding strength in silver oxide. Furthermore, in order to study the change in weight of the sample during decomposition of silver oxide, the TGA analysis was carried out. The TGA thermogram of control and treated samples is presented in Figure 3. The analysis result of TGA are presented in Table 4. The data exhibited that the control sample start losing the weight at onset temperature 335.0°C and stopped at endset temperature 407.0°C. However, in treated sample the onset and endset temperature were found at 322.45°C and 411.69°C respectively. Based on this data, the peak width *i.e.* the difference of onset and endset temperature, was found as 72°C in control sample and it was significantly increased to 89.24°C in treated silver oxide sample. In this decomposition process, the control and treated samples were lost 6.61 and 6.96% of its initial weight respectively. Also the rate of weight loss during decomposition was increased from 1.69×10^{-6} g/s (control) to 1.76×10^{-6} g/s in treated silver oxide sample. It indicated that rate of weight loss during decomposition of silver oxide was increased by 4.14%. The

increase in decomposition energy in DSC and rate of weight loss in TGA are contrary to each other. It is possible that the biofield energy altered the interatomic interaction in silver oxide, which resulted into change in the rate of weight loss during decomposition. It is reported that the change in

thermal characteristics can influenced the rate of dissolution of a drug [23]. Thus, it is assumed that the change in thermal characteristic of treated silver oxide might improve its rate of dissolution in formulation and thus bioavailability.

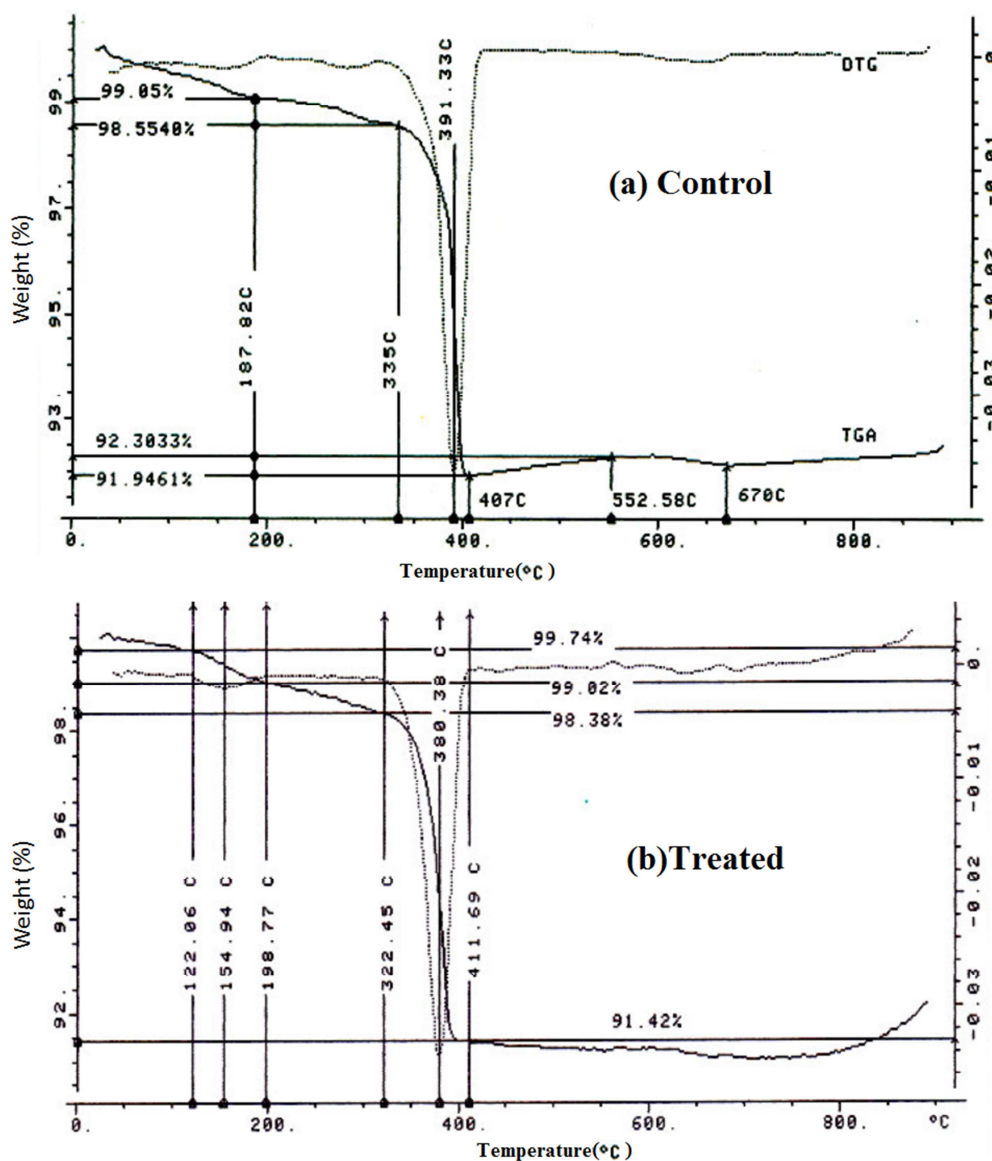


Fig. 3. TGA thermogram of silver oxide.

Table 4. TGA analysis of silver oxide powder.

Parameter	Control	T1
Onset temperature (°C)	335.00	322.45
Peak temperature (°C)	391.33	380.38
Endset temperature (°C)	407.00	411.69
Peak width (°C)	72.00	89.24
Weight of the sample at onset (%)	98.55	98.38
Weight of the sample at endset (%)	91.95	91.42
Change in weight (%)	-6.61	-6.96
Percent change in weight/width	-0.09	-0.08
Weight loss rate (g/s) $\times 10^{-6}$	-1.69	-1.76
Percent change in weight loss rate (%)	-	4.14

3.3. FT-IR Spectroscopy

The FT-IR spectra of control and treated silver oxide samples are presented in Figure 4. The control and treated samples showed the band at 3190 cm^{-1} , which was attributed to O-H stretching vibrations. It could be due to moisture absorption by the samples from the environment. In addition, the band observed at 1487 cm^{-1} in control and 1484 cm^{-1} in treated sample were due to C=O stretching vibrations [24]. It could be due to CO_2 absorption by the sample from the environment. The characteristic band of silver oxide bond was observed at 547 cm^{-1} in control and treated sample [25]. Thus, the FT-IR data did not show any significant alteration

in absorption wavenumbers of treated sample as compared to the control.

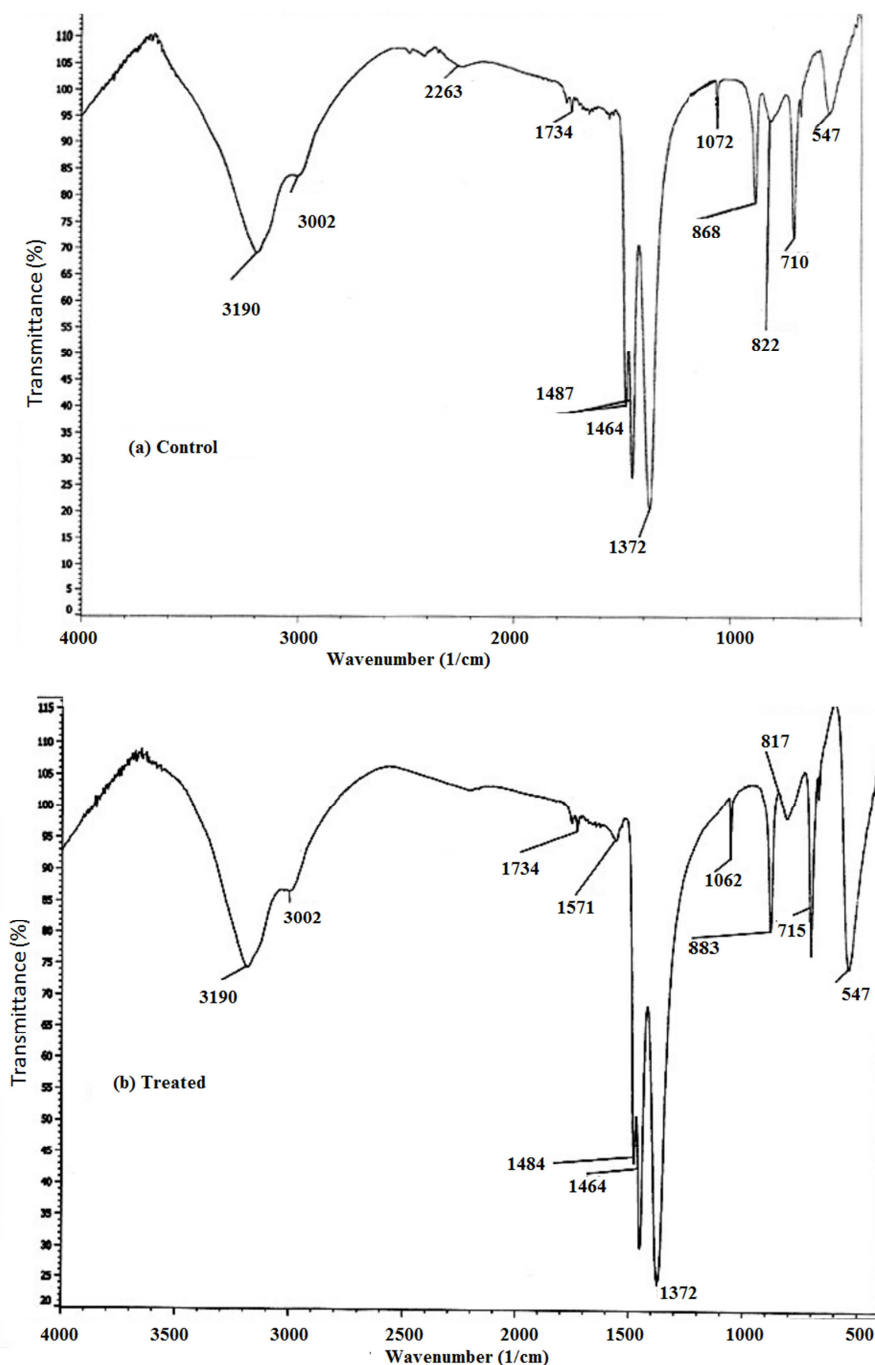


Fig. 4. FT-IR spectra of silver oxide powder.

4. Conclusions

The XRD data showed that the crystallite size was significantly increased upto 150% as compared to the control. The DSC data revealed that the thermal decomposition energy of silver oxide was increased up to 71.58% in treated sample as compared to the control. TGA data exhibited that the rate of weight loss due to decomposition of silver oxide was increased by 4.14% in treated sample as compared to the control. Hence, the

biofield treatment has significantly altered the physical and thermal properties of silver oxide. Based on the alteration, it is assumed that biofield energy treatment could be applied to silver oxide to improve its dissolution rate in topical formulation and bioavailability.

Acknowledgments

Authors thank Dr. Cheng Dong of NLSC, Institute of Physics, and Chinese academy of sciences for permitting us to use Powder-X software for analyzing XRD results. The

authors would also like to thank Trivedi Science, Trivedi Master Wellness and Trivedi Testimonials for their support during the work.

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