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Thermodynamic Studies of High-Temperature Processes in the System "Loesslike Loam - Coal - Alkaline"

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Abstract: In recent years, studies of a number of foreign scientists have been reported on the possibility of using loess loam as the main constituent of the ceramic composition in the production of clinker brick. Loess loam is considered a common raw material. In this paper, we consider the purpose, which is research on the study and development of methods that regulate the physicochemical processes occurring in the mass of loess-like loam under the influence of temperature. As a result of these studies it will be possible to intensify thermochemical processes and achieve high-quality products with minimal energy costs. Thus, the possible chemical interactions between minerals of loess-like loam have been theoretically studied, the medium of which is given priority to the decomposition of kaolinite. To intensify the decomposition process of refractory minerals, it is recommended to use an alkali additive in an amount of up to 5%, with a weight ratio of coal.

Keywords: Chemical Processes, Loess-Like Loam, Thermochemical Processes, Alkali Oxides, Chemical Reactions, Coke-Alkaline Additive, Carbon, Sodium Hydroxide

1. Introduction

Rapid development of construction of housing and industrial buildings and structures dictates the need to increase the output of wall materials. The wall building materials include ceramic bricks, obtained by burning in a high-temperature environment. The main cheap raw material for the production of ceramic building bricks is loess loam, whose reserves in Uzbekistan are unlimited. Loesslike loams in Uzbekistan were formed as a result of the deposition of dust brought by the wind mainly from deserts, and also there are silty deposits of various rivers carrying products of rock destruction in their waters. Under the influence of atmospheric factors, loose materials gradually disintegrated, enriching loam with finely dispersed fractions. The mineralogical composition of loesslike loams is represented mainly by quartz, carbonates, feldspars, biotite, muscovite and clay minerals. From the point of view of chemistry, loess-like loam is a ready-made mixture consisting of several active and inert components, among which there are basic, alkaline earth, alkaline oxides, as well as elements of variable valence. Under the influence of heat, there is an interaction between these substances, resulting in a stone-like body [1-5].

The purpose of our research is to study and develop methods that regulate the physicochemical processes occurring in the mass of loess-like loam under the influence of temperature. As a result of these studies it will be possible to intensify thermochemical processes and achieve high-quality products with minimal energy costs.

The issues of intensification of high-temperature processes and prediction of the properties of ceramic materials in the firing process at given temperatures and holding times are a necessary task in the light of solutions to energy conservation issues. In scientific and technical and applied research, thermodynamic analysis is often used to create and select optimal technological schemes for processing raw materials, as well as for preliminary modeling of technical processes and synthesis of materials with specified physical and technical properties. The thermodynamic method of investigating high-temperature processes can provide valuable information on the possibility of purposeful behavior of various reactions, on the ways of controlling

processes, on the thermal changes accompanying the reaction [6-7].

2. Methods

For chemical reactions taking place at P = const and T = const, according to the second law of thermodynamics, by means of which the energy possibility is determined, and the direction of the chemical reactions is expressed by the equation

$$\Delta G_T^0 = \Delta H_T^0 - T \cdot \Delta S_T^0$$

where, T is the absolute temperature ^oK, ΔH_T^0; $\cdot \Delta S_T^0$ - change in the thermal effect and entropy of the system under standard conditions (T = 25°C and P = 1 atm), respectively.

It is known that the thermal effect of the reaction is determined by a change in temperature in accordance with the Kirchhoff equation:

$$\Delta H_T^0 = \Delta H_{298}^0 + \int_{298}^{T} \Delta C_p \cdot dT$$

where,

ΔH_{298}^0 - the thermal effect of the reaction at standard temperature, ΔC_p is the difference between the total equilibrium constants, by means of which the equilibrium composition of the reacting mixture can be calculated, the limiting yield of products, the direction of the reaction is determined by the equation:

$$\lg K_p = \frac{-G_T^0}{2.3 \cdot RT}$$

Where

R is the universal gas constant,
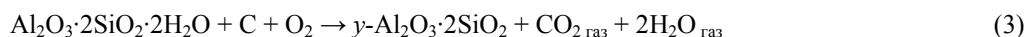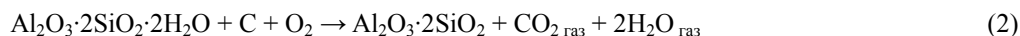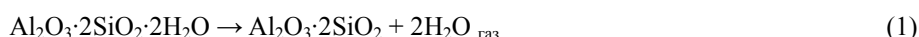G_T^0 is the standard Gibbs energy.

The solution of the main problems of thermodynamic investigation of high-temperature processes in solid mixtures requires careful consideration of phase transformations of components in the temperature range under consideration and changes in entropy and heat capacity with temperature. [8-10]

3. Materials

When the system goes into a more disorderly state, the entropy naturally increases. Therefore, the term of the equation $T \cdot \Delta S_T^0$, containing the entropy, always acts in such a way that as the temperature increases, the reaction proceeds in the direction associated with the increase in the amount of the liquid or gaseous phase. This explains, for example, that burning of carbon at high temperatures proceeds mainly by the equation, under any temperature conditions the reaction is accompanied by a thermal effect. This must be taken into account in the thermodynamic calculations of the reaction taking place at elevated temperatures involving carbon [11-14].

Main part. Under normal conditions, when fossil loam is fired in the temperature range 500-700°C, dehydration occurs and the structure of minerals is destroyed. In our studies, the purpose of intensifying the firing process of raw mixtures for the production of wall ceramic bricks was modified with coke-base and sodium hydroxide. In the case of the application of a mass of simultaneously burning additive and a low-melting component, i.e. alkali, the situation is different [2]. For the purpose of decomposition of sulphate, carbonate and other salts, burnt additives play the role of a reducing agent. Some reducing agents burn out at lower temperatures than the decomposition of salts. For example, sulfates decompose at higher (900°C) temperatures. To intensify the decomposition of refractory salts, a corresponding burn-out additive is necessary, the coke additive proposed is an appropriate component for these purposes [3]. For the energy evaluation of the use of this additive, it is necessary to evaluate the influence of each of its components. Therefore, the following reaction equations were computed for the calculation of ΔH_T^0, $\Delta G^o_{T,}$ K_p

$$Al_2O_3 \cdot 2SiO_2 \cdot 2H_2O \rightarrow Al_2O_3 \cdot 2SiO_2 + 2H_2O_{\text{газ}} \tag{1}$$

$$Al_2O_3 \cdot 2SiO_2 \cdot 2H_2O + C + O_2 \rightarrow Al_2O_3 \cdot 2SiO_2 + CO_{2\,\text{газ}} + 2H_2O_{\text{газ}} \tag{2}$$

$$Al_2O_3 \cdot 2SiO_2 \cdot 2H_2O + C + O_2 \rightarrow y\text{-}Al_2O_3 \cdot 2SiO_2 + CO_{2\,\text{газ}} + 2H_2O_{\text{газ}} \tag{3}$$

In the reaction equations, the percentage of carbon and NaOH in the coke-alkaline additive was taken into account, with its optimal introduction in relation to the clay component of the brick raw meal.

Using the thermodynamic properties of compounds, a calculation was made of the determination ΔH_T^0 f the thermal effect of the reaction (1)-(3) ΔG_T^0 of the probability of flow, K_p of the equilibrium constant in the temperature range 473-723 K (200-450°C).

4. Results

The thermal effect of reaction (1)-(6) with temperature changes can be calculated by the equations:

$$\Delta H_{T_1}^0 = 342,80 \cdot 10^3 + \int_{298}^{T} (49,11 - 89,44 \cdot 10^{-3} T + 18,97 \cdot 10^5 \cdot T^{-2}) dT$$

$$\Delta H_{T_2}^0 = -50,71 \cdot 10^3 + \int_{298}^{T} (44,29 - 88,06 \cdot 10^{-3} T + 23,11 \cdot 10^5 \cdot T^{-2}) dT$$

$$\Delta H_{T_3}^0 = -237{,}96 \cdot 10^3 + \int_{298}^{T}(-22{,}61 - 9{,}09 \cdot 10^{-3}T + 13{,}79 \cdot 10^5 \cdot T^{-2})dT$$

The equations ΔS_T for the reaction (1)-(3) have the following form:

$$\Delta S_T = \Delta S_{298} + \int_{298}^{T}\frac{C_P}{T} \cdot dT$$

$$\Delta S_{T_1} = 332{,}19 + \int_{298}^{T}\frac{(49{,}11 - 89{,}44 \cdot 10^{-3}T + 18{,}97 \cdot 10^5 \cdot T^{-2})}{T} \cdot dT$$

$$\Delta S_{T_2} = 335{,}02 + \int_{298}^{T}\frac{(44{,}29 - 88{,}06 \cdot 10^{-3}T + 23{,}11 \cdot 10^5 \cdot T^{-2})}{T} \cdot dT$$

$$\Delta S_{T_3} = 335{,}02 + \int_{298}^{T}\frac{(-22{,}61 - 9{,}09 \cdot 10^{-3}T + 13{,}79 \cdot 10^5 \cdot T^{-2})}{T} \cdot dT$$

Table 1. *The value of ΔS_T^0, J/mol, ΔG_T^0, kJ/mol, K_p as a function of temperature T.*

Reaction	Temperature T, K		
	473		
	ΔS_T^0	ΔG_T^0	K_p
1	326,3	193,35	$4{,}5 \cdot 10^{-22}$
2	324,7	-199,6	$1{,}1 \cdot 10^{22}$
3	191,7	-322,8	$4{,}4 \cdot 10^{35}$

Table 2. *The value of ΔS_T^0, J/mol, ΔG_T^0, kJ/mol, K_p as a function of temperature T.*

Reaction	Temperature T, K		
	600		
	ΔS_T^0	ΔG_T^0	K_p
1	323,5	154,61	$3{,}5 \cdot 10^{-14}$
2	319,8	-205,1	$2{,}1 \cdot 10^{21}$
3	195,9	-351,7	$4{,}1 \cdot 10^{30}$

Table 3. *The value of ΔS_T^0, J/mol, ΔG_T^0, kJ/mol, K_p as a function of temperature T.*

Reaction	Temperature T, K		
	725		
	ΔS_T^0	ΔG_T^0	K_p
1	319,9	116,05	$4{,}4 \cdot 10^{-9}$
2	314,7	-236,2	$1{,}2 \cdot 10^{17}$
3	202,8	-379,5	$2{,}2 \cdot 10^{27}$

5. Discussion

The results of the calculations ΔH_T^0 indicate that the reaction (1) dehydration of kaolinite without additives is an endothermic process. The introduction of an additive containing carbon, the same reaction (2, 3) dehydration already takes place with the release of heat $\Delta H_T^0 < 0$,, i.e carbon in the presence of oxygen burns, and this heat is expended on the destruction of the crystal lattice of minerals. The amount of heat released with a sufficient amount of carbon is for the reaction (2) 46.03-79.88 kJ / mol. The presence of alkalis intensifies many high-temperature physico-chemical synthesis processes, including the dehydration of clay minerals, and with increasing temperature, the amount of heat liberated increases and reaches 84.48 kJ / mol at 725°K. The introduction of a complex additive containing carbon and sodium hydroxide further accelerates the process of dehydration of kaolinite, the amount of heat released is in the temperature range 473-725°K 80.25-119 kJ / mol, and taking into account the ratio of carbon, NaOH, both separately and in the complex, it is possible to shift the equilibrium from left to right and the possibility of the dehydration reaction of kaolinite in the temperature range under consideration is present ($\Delta G_T^0 < 0$, a $K_p > 1$).

Moreover, as the temperature increases, ΔG_T^0 decrease and K_p increase, which indicates an increase in the stability and stability of the system [15-17].

6. Conclusion

Thus, the possible chemical interactions between minerals of loess-like loam have been theoretically studied, the medium of which is given priority to the decomposition of kaolinite. To intensify the decomposition process of refractory minerals, it is recommended to use an alkali additive in an amount of up to 5%, with a weight ratio of coal.

Such an approach not only improves the rheological properties of the raw composition for the production of high-quality bricks, but also allows the dehydration of clay minerals to be shifted to lower temperatures, and the amount of heat released in the temperature range 473-725°K will lead to significant savings in thermal energy.

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